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## Structure Reports

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# {4,4'-Dimethyl-2,2'-[2,2-dimethylpropane-1,3-diylbis(nitrilomethanylidene)]diphenolato}copper(II) monohydrate

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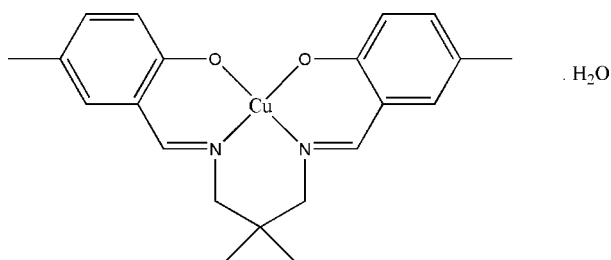
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Key indicators: single-crystal X-ray study;  $T = 296$  K; mean  $\sigma(\text{C}-\text{C}) = 0.012$  Å;  $R$  factor = 0.084;  $wR$  factor = 0.209; data-to-parameter ratio = 14.2.

The asymmetric unit of the title compound,  $[\text{Cu}(\text{C}_{21}\text{H}_{24}\text{N}_2\text{O}_2)] \cdot \text{H}_2\text{O}$ , comprises half of a Schiff base complex and half of a water molecule. The whole compound is generated by crystallographic twofold rotation symmetry. The geometry around the  $\text{Cu}^{\text{II}}$  atom, located on a twofold axis, is distorted square-planar, which is supported by the  $\text{N}_2\text{O}_2$  donor atoms of the coordinating Schiff base ligand. The dihedral angle between the symmetry-related benzene rings is  $47.5$  ( $4$ )°. In the crystal, the water molecule that is hydrogen bonded to the coordinated O atoms links the molecules *via*  $\text{O}-\text{H} \cdots \text{O}$  interactions into chains parallel to  $[001]$ . The crystal structure is further stabilized by  $\text{C}-\text{H} \cdots \pi$  interactions, and by  $\pi-\pi$  interactions involving inversion-related chelate rings [centroid-centroid distance =  $3.480$  ( $4$ ) Å].

## Related literature

For applications of Schiff bases in coordination chemistry, see: Granovski *et al.* (1993); Blower (1998). For related structures, see: Ghaemi *et al.* (2011); Kargar *et al.* (2011, 2012). For standard bond lengths, see: Allen *et al.* (1987).



## Experimental

### Crystal data

$[\text{Cu}(\text{C}_{21}\text{H}_{24}\text{N}_2\text{O}_2)] \cdot \text{H}_2\text{O}$   
 $M_r = 417.98$

Monoclinic,  $C2/c$  $a = 13.353$  (5) Å $b = 15.986$  (5) Å $c = 10.023$  (5) Å $\beta = 104.696$  (5)° $V = 2069.5$  (14) Å<sup>3</sup> $Z = 4$ Mo  $K\alpha$  radiation $\mu = 1.08$  mm<sup>-1</sup> $T = 296$  K $0.11 \times 0.08 \times 0.05$  mm

### Data collection

Bruker SMART APEXII CCD  
 area-detector diffractometer

Absorption correction: multi-scan  
 (SADABS; Bruker, 2005)

 $T_{\text{min}} = 0.891$ ,  $T_{\text{max}} = 0.948$ 

4967 measured reflections

1779 independent reflections

1053 reflections with  $I > 2\sigma(I)$  $R_{\text{int}} = 0.101$ 

### Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.084$  $wR(F^2) = 0.209$  $S = 0.95$ 

1779 reflections

125 parameters

H-atom parameters constrained

 $\Delta\rho_{\text{max}} = 1.03$  e Å<sup>-3</sup> $\Delta\rho_{\text{min}} = -0.96$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

Cg1 is the centroid of the C1–C6 ring.

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
O1W–H1W1 $\cdots$ O1	0.85	2.46	2.783 (7)	103
O1W–H1W1 $\cdots$ O1 <sup>i</sup>	0.85	2.44	2.783 (7)	105
C3–H3 $\cdots$ O1W <sup>ii</sup>	0.93	2.55	3.48 (1)	173
C8–H8B $\cdots$ Cg1 <sup>iii</sup>	0.97	2.83	3.693 (9)	148
C11–H11B $\cdots$ Cg1 <sup>iv</sup>	0.96	2.98	3.850 (12)	151

Symmetry codes: (i)  $-x + 1, y, -z - \frac{1}{2}$ ; (ii)  $-x + \frac{1}{2}, -y - \frac{1}{2}, -z - 1$ ; (iii)  $x + \frac{3}{2}, y + \frac{1}{2}, z - 1$ ; (iv)  $-x + \frac{1}{2}, y + \frac{1}{2}, -z - \frac{1}{2}$ .

Data collection: APEX2 (Bruker, 2005); cell refinement: SAINT (Bruker, 2005); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL and PLATON (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SU2488).

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## supporting information

*Acta Cryst.* (2012). E68, m1172 [doi:10.1107/S1600536812034502]

## {4,4'-Dimethyl-2,2'-[2,2-dimethylpropane-1,3-diylbis(nitrilomethanylylidene)]diphenolato}copper(II) monohydrate

Hadi Kargar, Reza Kia, Fatemeh Ganji and Valiollah Mirkhani

### S1. Comment

Schiff base complexes are one of the most important stereochemical models in transition metal coordination chemistry, with the ease of preparation and structural variations (Granovski *et al.*, 1993; Blower, 1998). In continuation of our work on the structural analysis of Schiff base metal complexes (Kargar *et al.*, 2012; Kargar *et al.*, 2011; Ghaemi, *et al.*, (2011), we synthesized the title compound and report herein on its crystal structure.

The asymmetric unit of the title compound, Fig. 1, comprises half of a Schiff base complex and half a water molecule. The Cu1 and C9 atoms of the complex and the O atom of the water molecule lie on a two-fold rotation axis which generates the whole complex. The bond lengths (Allen *et al.*, 1987) and angles are within the normal ranges and are comparable to those reported for related structures (Kargar *et al.*, 2012; Kargar *et al.*, 2011; Ghaemi *et al.*, (2011). The geometry around the Cu<sup>II</sup> atom is distorted square-planar which is supported by the N<sub>2</sub>O<sub>2</sub> donor atoms of the coordinated Schiff base ligand. The dihedral angle between the substituted benzene rings is 47.5 (4)°.

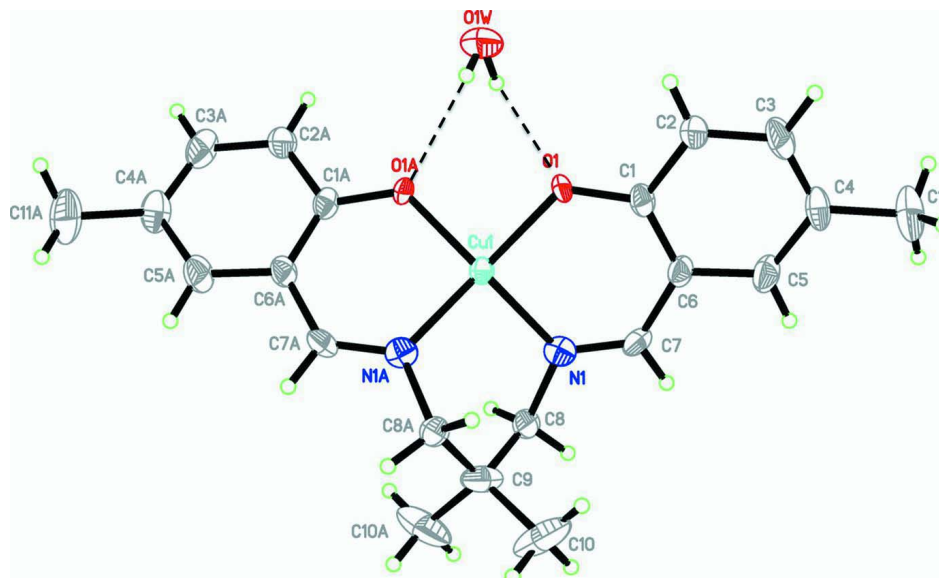
In the crystal, the water molecule that is hydrogen bonded to the coordinated O atoms, O1, mediates linking of molecules by C—H...O interactions (Table 1 and Fig. 2). The crystal structure is further stabilized by C-H... $\pi$  interactions (Table 1), and by  $\pi$ - $\pi$  interactions involving inversion related chelate rings [ $Cg \cdots Cg^i = 3.480$  (4) Å; Cg is the centroid of the Cu1/O1/C1/C6/C7/N1 ring; symmetry code: (i) 1 - x, -y, -1 - z].

### S2. Experimental

The title compound was synthesized by adding 5-methyl-salicylaldehyde-2,2-dimethyl-1,3-propanediamine (2 mmol) to a solution of CuCl<sub>2</sub>·4H<sub>2</sub>O (2.1 mmol) in ethanol (30 ml). The mixture was refluxed with stirring for 30 min. The resultant solution was filtered. Dark-green single crystals of the title compound suitable for X-ray structure determination were recrystallized from ethanol by slow evaporation of the solvents at room temperature over several days.

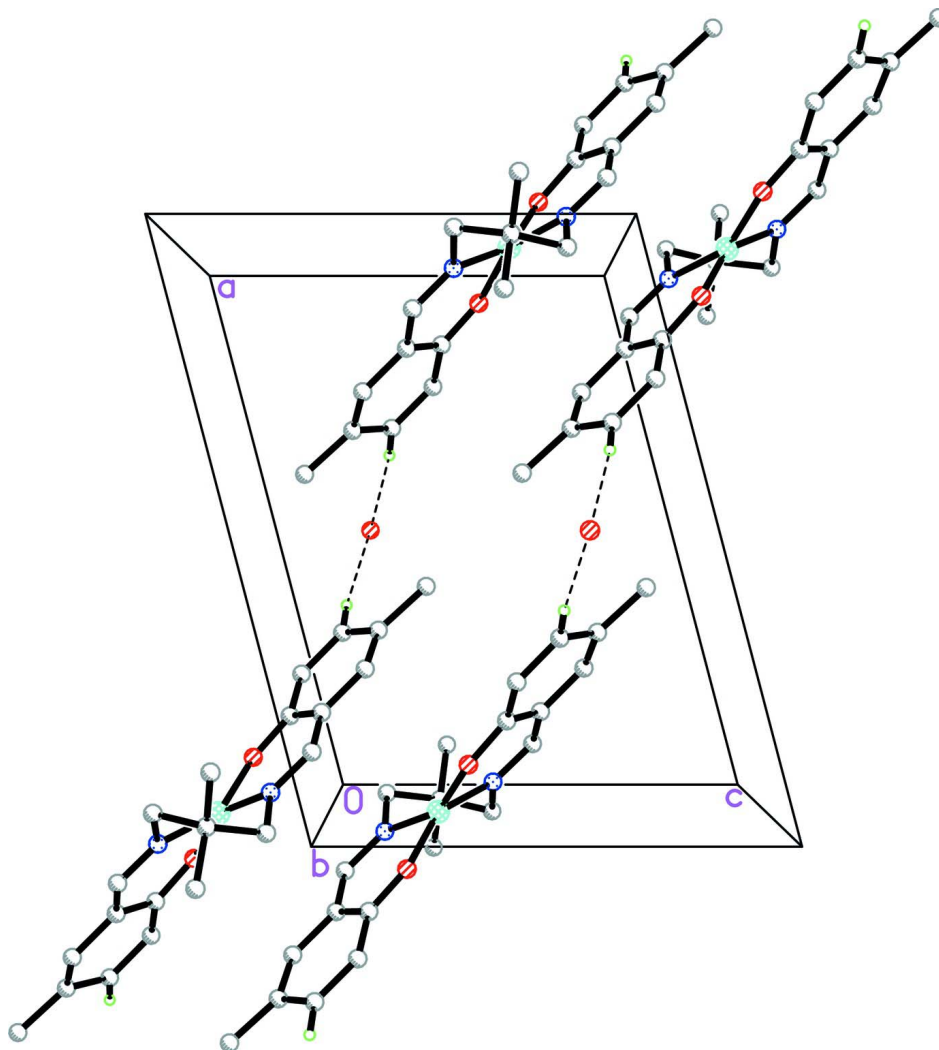
### S3. Refinement

The water H atom was located in a difference Fourier map and refined as a riding atom with  $U_{iso}(H) = 1.5U_{eq}(O)$ . The C-bound H-atoms were included in calculated positions and treated as riding atoms: C—H = 0.93, 0.96 and 0.97 Å for CH, CH<sub>3</sub> and CH<sub>2</sub> H-atoms, respectively, with  $U_{iso}(H) = k \times U_{eq}(C)$ , where  $k = 1.5$  for CH<sub>3</sub> H-atoms, and = 1.2 for other H-atoms.



**Figure 1**

A view of the molecular structure of the title compound, with the atom numbering. The displacement ellipsoids are drawn at the 40% probability level. The O-H...O hydrogen bonds are shown as dashed lines (see Table 1 for details; symmetry code for suffix A =  $-x+1, y, -z-1/2$ ).



**Figure 2**

A view along the *b* axis of the crystal packing of the title compound showing the C—H...O interactions as dashed lines [see Table 1 for details; only the H atoms involved in these interactions are shown].

**{4,4'-Dimethyl-2,2'-[2,2-dimethylpropane-1,3-diylbis(nitrilomethanylylidene)]diphenolato}copper(II) monohydrate**

*Crystal data*

[Cu(C<sub>21</sub>H<sub>24</sub>N<sub>2</sub>O<sub>2</sub>)]·H<sub>2</sub>O

*M<sub>r</sub>* = 417.98

Monoclinic, *C*2/*c*

Hall symbol: -*C* 2yc

*a* = 13.353 (5) Å

*b* = 15.986 (5) Å

*c* = 10.023 (5) Å

β = 104.696 (5)°

*V* = 2069.5 (14) Å<sup>3</sup>

*Z* = 4

*F*(000) = 876

*D<sub>x</sub>* = 1.342 Mg m<sup>-3</sup>

Mo *K*α radiation, λ = 0.71073 Å

Cell parameters from 512 reflections

θ = 2.5–27.4°

μ = 1.08 mm<sup>-1</sup>

*T* = 296 K

Block, dark-green

0.11 × 0.08 × 0.05 mm

*Data collection*

Bruker SMART APEXII CCD area-detector  
diffractometer  
Radiation source: fine-focus sealed tube  
Graphite monochromator  
 $\varphi$  and  $\omega$  scans  
Absorption correction: multi-scan  
(*SADABS*; Bruker, 2005)  
 $T_{\min} = 0.891$ ,  $T_{\max} = 0.948$

4967 measured reflections  
1779 independent reflections  
1053 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.101$   
 $\theta_{\max} = 25.0^\circ$ ,  $\theta_{\min} = 2.0^\circ$   
 $h = -15 \rightarrow 11$   
 $k = -18 \rightarrow 18$   
 $l = -10 \rightarrow 11$

*Refinement*

Refinement on  $F^2$   
Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.084$   
 $wR(F^2) = 0.209$   
 $S = 0.95$   
1779 reflections  
125 parameters  
0 restraints  
Primary atom site location: structure-invariant  
direct methods

Secondary atom site location: difference Fourier  
map  
Hydrogen site location: inferred from  
neighbouring sites  
H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.0775P)^2]$   
where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} < 0.001$   
 $\Delta\rho_{\max} = 1.03 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.96 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted R-factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional R-factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\sigma(F^2)$  is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and R-factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cu1	0.5000	0.01026 (6)	-0.2500	0.0266 (5)
O1	0.4089 (4)	-0.0737 (2)	-0.3499 (6)	0.0312 (14)
N1	0.4554 (5)	0.0944 (3)	-0.3912 (7)	0.0305 (16)
C1	0.3331 (6)	-0.0612 (4)	-0.4620 (9)	0.0293 (19)
C2	0.2617 (6)	-0.1276 (4)	-0.5073 (10)	0.042 (3)
H2	0.2662	-0.1758	-0.4541	0.050*
C3	0.1860 (7)	-0.1214 (5)	-0.6291 (11)	0.046 (3)
H3	0.1375	-0.1642	-0.6530	0.055*
C4	0.1790 (6)	-0.0526 (5)	-0.7191 (10)	0.045 (2)
C6	0.3222 (6)	0.0109 (4)	-0.5441 (9)	0.0311 (19)
C7	0.3854 (6)	0.0845 (4)	-0.5067 (10)	0.032 (2)
H7	0.3753	0.1282	-0.5699	0.038*
C8	0.5166 (6)	0.1727 (4)	-0.3691 (9)	0.036 (2)
H8A	0.4984	0.2056	-0.4531	0.043*
H8B	0.5894	0.1587	-0.3512	0.043*
C9	0.5000	0.2260 (6)	-0.2500	0.048 (4)

C5	0.2449 (6)	0.0134 (5)	-0.6699 (10)	0.042 (2)
H5	0.2380	0.0621	-0.7224	0.051*
C11	0.1004 (8)	-0.0471 (7)	-0.8564 (12)	0.069 (3)
H11B	0.1206	-0.0042	-0.9113	0.104*
H11A	0.0338	-0.0338	-0.8421	0.104*
H11C	0.0968	-0.0999	-0.9032	0.104*
C10	0.4033 (10)	0.2806 (6)	-0.2988 (13)	0.087 (5)
H10B	0.3947	0.3149	-0.2238	0.130*
H10A	0.3436	0.2455	-0.3299	0.130*
H10C	0.4110	0.3157	-0.3733	0.130*
O1W	0.5000	-0.2254 (4)	-0.2500	0.091 (5)
H1W1	0.5281	-0.1936	-0.2981	0.137*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cu1	0.0194 (7)	0.0263 (6)	0.0300 (9)	0.000	-0.0015 (6)	0.000
O1	0.026 (3)	0.026 (2)	0.031 (4)	-0.0008 (18)	-0.012 (3)	-0.005 (2)
N1	0.033 (4)	0.032 (3)	0.030 (5)	-0.005 (2)	0.016 (4)	0.003 (3)
C1	0.018 (4)	0.038 (4)	0.030 (5)	0.001 (3)	0.003 (4)	-0.005 (3)
C2	0.031 (5)	0.037 (4)	0.049 (7)	0.000 (3)	-0.007 (5)	0.000 (4)
C3	0.020 (5)	0.060 (5)	0.052 (7)	-0.007 (3)	0.001 (5)	-0.014 (5)
C4	0.017 (4)	0.068 (5)	0.046 (7)	-0.005 (4)	-0.003 (5)	-0.003 (5)
C6	0.022 (4)	0.036 (3)	0.032 (5)	0.007 (3)	0.001 (4)	0.001 (3)
C7	0.030 (5)	0.029 (3)	0.036 (6)	0.007 (3)	0.006 (5)	0.011 (3)
C8	0.026 (5)	0.032 (4)	0.044 (6)	-0.007 (3)	-0.002 (5)	0.004 (3)
C9	0.046 (9)	0.025 (5)	0.073 (12)	0.000	0.018 (8)	0.000
C5	0.026 (5)	0.052 (4)	0.043 (6)	0.005 (3)	-0.002 (4)	0.007 (4)
C11	0.036 (6)	0.113 (8)	0.049 (8)	-0.013 (5)	-0.007 (6)	0.000 (6)
C10	0.123 (12)	0.058 (6)	0.089 (11)	0.050 (6)	0.045 (9)	0.034 (6)
O1W	0.106 (10)	0.033 (5)	0.106 (11)	0.000	-0.026 (8)	0.000

*Geometric parameters (Å, °)*

Cu1—O1 <sup>i</sup>	1.914 (4)	C6—C7	1.441 (10)
Cu1—O1	1.914 (4)	C7—H7	0.9300
Cu1—N1 <sup>i</sup>	1.934 (6)	C8—C9	1.527 (10)
Cu1—N1	1.934 (6)	C8—H8A	0.9700
O1—C1	1.322 (9)	C8—H8B	0.9700
N1—C7	1.300 (10)	C9—C8 <sup>i</sup>	1.527 (10)
N1—C8	1.480 (8)	C9—C10	1.533 (10)
C1—C6	1.403 (10)	C9—C10 <sup>i</sup>	1.533 (10)
C1—C2	1.423 (10)	C5—H5	0.9300
C2—C3	1.377 (12)	C11—H11B	0.9600
C2—H2	0.9300	C11—H11A	0.9600
C3—C4	1.411 (13)	C11—H11C	0.9600
C3—H3	0.9300	C10—H10B	0.9600
C4—C5	1.382 (11)	C10—H10A	0.9600

C4—C11	1.507 (12)	C10—H10C	0.9600
C6—C5	1.413 (11)	O1W—H1W1	0.8513
O1 <sup>i</sup> —Cu1—O1	91.0 (3)	N1—C8—C9	113.9 (7)
O1 <sup>i</sup> —Cu1—N1 <sup>i</sup>	93.9 (2)	N1—C8—H8A	108.8
O1—Cu1—N1 <sup>i</sup>	155.2 (3)	C9—C8—H8A	108.8
O1 <sup>i</sup> —Cu1—N1	155.1 (3)	N1—C8—H8B	108.8
O1—Cu1—N1	93.9 (2)	C9—C8—H8B	108.8
N1 <sup>i</sup> —Cu1—N1	91.9 (4)	H8A—C8—H8B	107.7
C1—O1—Cu1	125.9 (4)	C8—C9—C8 <sup>i</sup>	112.2 (8)
C7—N1—C8	118.8 (6)	C8—C9—C10	110.2 (6)
C7—N1—Cu1	125.8 (4)	C8 <sup>i</sup> —C9—C10	106.9 (6)
C8—N1—Cu1	115.0 (5)	C8—C9—C10 <sup>i</sup>	106.9 (6)
O1—C1—C6	124.5 (6)	C8 <sup>i</sup> —C9—C10 <sup>i</sup>	110.2 (6)
O1—C1—C2	117.8 (7)	C10—C9—C10 <sup>i</sup>	110.5 (11)
C6—C1—C2	117.6 (8)	C4—C5—C6	123.4 (8)
C3—C2—C1	120.8 (8)	C4—C5—H5	118.3
C3—C2—H2	119.6	C6—C5—H5	118.3
C1—C2—H2	119.6	C4—C11—H11B	109.5
C2—C3—C4	122.5 (7)	C4—C11—H11A	109.5
C2—C3—H3	118.8	H11B—C11—H11A	109.5
C4—C3—H3	118.8	C4—C11—H11C	109.5
C5—C4—C3	115.8 (8)	H11B—C11—H11C	109.5
C5—C4—C11	121.0 (9)	H11A—C11—H11C	109.5
C3—C4—C11	123.1 (8)	C9—C10—H10B	109.5
C1—C6—C5	119.4 (7)	C9—C10—H10A	109.5
C1—C6—C7	123.4 (7)	H10B—C10—H10A	109.5
C5—C6—C7	117.1 (7)	C9—C10—H10C	109.5
N1—C7—C6	124.9 (7)	H10B—C10—H10C	109.5
N1—C7—H7	117.5	H10A—C10—H10C	109.5
C6—C7—H7	117.5		
O1 <sup>i</sup> —Cu1—O1—C1	-166.5 (8)	C2—C1—C6—C5	-2.8 (12)
N1 <sup>i</sup> —Cu1—O1—C1	92.1 (9)	O1—C1—C6—C7	-7.9 (13)
N1—Cu1—O1—C1	-10.9 (7)	C2—C1—C6—C7	176.4 (8)
O1 <sup>i</sup> —Cu1—N1—C7	101.3 (8)	C8—N1—C7—C6	178.2 (8)
O1—Cu1—N1—C7	0.5 (8)	Cu1—N1—C7—C6	5.8 (13)
N1 <sup>i</sup> —Cu1—N1—C7	-155.3 (9)	C1—C6—C7—N1	-3.6 (14)
O1 <sup>i</sup> —Cu1—N1—C8	-71.3 (8)	C5—C6—C7—N1	175.7 (8)
O1—Cu1—N1—C8	-172.1 (6)	C7—N1—C8—C9	116.3 (8)
N1 <sup>i</sup> —Cu1—N1—C8	32.1 (4)	Cu1—N1—C8—C9	-70.5 (7)
Cu1—O1—C1—C6	15.7 (11)	N1—C8—C9—C8 <sup>i</sup>	35.3 (4)
Cu1—O1—C1—C2	-168.6 (6)	N1—C8—C9—C10	-83.7 (9)
O1—C1—C2—C3	-174.8 (8)	N1—C8—C9—C10 <sup>i</sup>	156.2 (8)
C6—C1—C2—C3	1.2 (14)	C3—C4—C5—C6	6.1 (15)
C1—C2—C3—C4	4.2 (15)	C11—C4—C5—C6	-177.9 (9)
C2—C3—C4—C5	-7.7 (15)	C1—C6—C5—C4	-1.0 (14)

C2—C3—C4—C11	176.4 (10)	C7—C6—C5—C4	179.7 (9)
O1—C1—C6—C5	172.9 (8)		

Symmetry code: (i)  $-x+1, y, -z-1/2$ .

*Hydrogen-bond geometry* ( $\text{\AA}$ ,  $^\circ$ )

Cg1 is the centroid of the C1—C6 ring.

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
O1 <i>W</i> —H1 <i>W</i> 1 $\cdots$ O1	0.85	2.46	2.783 (7)	103
O1 <i>W</i> —H1 <i>W</i> 1 $\cdots$ O1 <sup>i</sup>	0.85	2.44	2.783 (7)	105
C3—H3 $\cdots$ O1 <i>W</i> <sup>ii</sup>	0.93	2.55	3.48 (1)	173
C8—H8 <i>B</i> $\cdots$ Cg1 <sup>iii</sup>	0.97	2.83	3.693 (9)	148
C11—H11 <i>B</i> $\cdots$ Cg1 <sup>iv</sup>	0.96	2.98	3.850 (12)	151

Symmetry codes: (i)  $-x+1, y, -z-1/2$ ; (ii)  $-x+1/2, -y-1/2, -z-1$ ; (iii)  $x+3/2, y+1/2, z-1$ ; (iv)  $-x+1/2, y+1/2, -z-1/2$ .