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1,4-Dihydrobenzo[g]quinoxaline-2,3-dione

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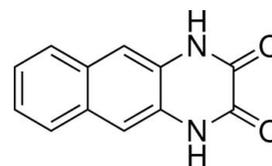
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.094; wR factor = 0.271; data-to-parameter ratio = 11.6.

The title compound, $\text{C}_{12}\text{H}_8\text{N}_2\text{O}_2$, was prepared by the reaction of the diethyl ester of naphthalenebis(oxamate) with *tert*-BuNH₂. The molecule is nearly planar, with an r.m.s. deviation of 0.017 Å from the plane through all 16 non-H atoms. In the crystal, a three-dimensional network is formed, composed of layers of molecules along the *b*- and *c*-axis directions, due to the formation of intermolecular N—H···O hydrogen bonds, as well as of chains along the *a*-axis direction due to parallel displaced sandwich-type π – π interactions with average distances between the interacting molecules in the range 3.35–3.40 Å.

Related literature

For the synthesis and structure of 1,4-dihydrobenzo[g]quinoxaline-2,3-dione·3H₂O, see: Oxtoby *et al.* (2005). For the use of bis(oxamates) and bis(oxamidates) for complex formation, see: Pardo *et al.* (2008) and Abdulmalic *et al.* (2012); Rüffer *et al.* (2012), respectively. For the general synthesis of bis(oxamidates), see: Ruiz *et al.* (1997) and for the synthesis of diethyl *N,N'*-naphthalene-bis(oxamate), see: Rüffer *et al.* (2007). For thin film formation by bis(oxamato) complexes, see: Eya'ane Meva (2009); Bräuer *et al.* (2006). For self-organization in supramolecular chemistry due to intermolecular π interactions and/or hydrogen bonds, see: Burrow *et al.* (1996); Chowdhry *et al.* (1996); Dai *et al.* (1997); Munoz *et al.* (1998). For dione tautomerism in the solid state, see: Svenson (1976).



Experimental

Crystal data

$\text{C}_{12}\text{H}_8\text{N}_2\text{O}_2$
 $M_r = 212.20$
 Monoclinic, $P2_1/n$
 $a = 7.1334$ (15) Å
 $b = 8.4229$ (18) Å
 $c = 15.292$ (2) Å
 $\beta = 99.792$ (14)°

$V = 905.4$ (3) Å³
 $Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 0.11$ mm⁻¹
 $T = 293$ K
 $0.3 \times 0.2 \times 0.1$ mm

Data collection

Oxford Gemini S diffractometer
 Absorption correction: multi-scan
 (*CrysAlis RED*; Oxford
 Diffraction, 2006)
 $T_{\min} = 0.649$, $T_{\max} = 1.000$

5294 measured reflections
 1773 independent reflections
 875 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.039$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.094$
 $wR(F^2) = 0.271$
 $S = 0.92$
 1773 reflections
 153 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\max} = 0.77$ e Å⁻³
 $\Delta\rho_{\min} = -0.57$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> —H··· <i>A</i>	<i>D</i> —H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> —H··· <i>A</i>
N1—H1N···O1 ¹	0.91 (4)	2.27 (4)	2.866 (3)	122 (3)
N2—H2N···O2 ²ⁱ	0.94 (4)	1.90 (4)	2.843 (4)	176 (4)

Symmetry codes: (i) $-x + \frac{1}{2}, y + \frac{1}{2}, -z + \frac{3}{2}$; (ii) $-x, -y - 1, -z + 2$.

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2006); cell refinement: *CrysAlis CCD*; data reduction: *CrysAlis RED* (Oxford Diffraction, 2006); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 2012); software used to prepare material for publication: *WinGX* (Farrugia, 2012) and *publCIF* (Westrip, 2010).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SJ5284).

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supporting information

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1,4-Dihydrobenzo[g]quinoxaline-2,3-dione

François Eya'ane Meva, Dieter Schaarschmidt, Mohammad A. Abdulmalic and Tobias Rüffer

S1. Comment

Weak interactions such as intermolecular π interactions or the formation of hydrogen bonds are essential for molecular recognition and self organization in biological systems and supramolecular chemistry (Burrow *et al.*, 1996; Chowdhry *et al.*, 1996; Dai *et al.*, 1997). Oxtoby *et al.* (2005) synthesized derivatives of dihydroquinoxalinedione containing a hydrophilic oxalamide-based "terminus". This oxalamide serves to increase the water solubility of the organic molecules and allows H₂O molecules to be hydrogen bonded to the organic molecules as well as to each other. In this perspective, 1,4-dihydrobenzo[g]quinoxaline-2,3-dione hydrate (**1**×3H₂O) has been crystallized by slowly cooling a boiling solution of the powder in DMF/water to room temperature (Oxtoby *et al.*, 2005). Infinite arrays containing two different alternating head-to-tail π interactions parallel to the crystallographic *a* axis were observed in the solid state of **1**×3H₂O, with the π stacks being orthogonal to chains of H₂O molecules and held together by $R^2_2(8)$ hydrogen-bonding interactions.

Bis(oxamate) molecules (type **I** molecules, *cf.* Figure 4) have been widely used in order to produce a series of complexes with various magnetic interactions (Pardo *et al.*, 2008). In previous work, our interest was devoted to a study of the impact of π interactions in the formation of thin films using trimetallic bis(oxamato) type complexes (Eya'ane Meva, 2009; Bräuer *et al.*, 2006) as well as the investigation of electronic effects on magnetic *J* couplings (Rüffer *et al.*, 2007). With the aim of using type **I**—NR₂ molecules as ligands for the formation of transition metal complexes (Abdulmalic *et al.*, 2012; Rüffer *et al.*, 2012), we became interested in their synthesis. Generally, type **I** molecules are reacted with an excess of a primary amine (Ruiz *et al.*, 1997) resulting in the formation of type **I**—NR₂ molecules. On the other hand, quinoxaline derivatives are generally synthesized by refluxing diamines and oxalic acid in HCl (Oxtoby *et al.*, 2005). We report here, that the reaction of the diethyl ester of naphthalene-bis(oxamate) (Rüffer *et al.*, 2007), Fig. 4, with an excess of *tert*-BuNH₂ in MeOH does not give the corresponding type **I**—NR₂ molecule, but instead forms 1,4-dihydrobenzo[g]quinoxaline-2,3-dione (**1**). A similar reaction has been already described by Munoz *et al.* (1998), who treated *o*-phenylenebis(oxamate) with [Me₄N]OH and Fe(ClO₄)₃ to obtain an analogous derivatized product in the form of its Fe(III) complex.

In the solid state **1** is nearly planar (r. m. s. d. of a calculated mean plane of C1–C12, O1, O2, N1, N2 = 0.017 Å). The bond lengths of the C=O functions of **1** (1.230 (4) and 1.240 (4) Å) reveal the molecule to be the dione tautomer in the solid state (Svenson, 1976). All other bond distances and angles are comparable those described for **1**×3H₂O (Oxtoby *et al.*, 2005).

In the crystal structure of **1** the formation of a three-dimensional network is observed. Intermolecular hydrogen bonds between individual molecules of **1**, *cf.* Table 1, form two-dimensional layers. A representative view of one selected two-dimensional layer is illustrated in Figure 2, showing the two-dimensional layers extending along the crystallographic *b*- and *c*-axes. Within the two-dimensional layers the formation of dimers of **1** with $R^2_2(8)$ type hydrogen bond interactions is also observed, as reported for **1**×3H₂O (Oxtoby *et al.*, 2005), *cf.* N2—H2N[⋯]O2ⁱⁱ in Table 1 and Figure 2. However, due to the N1—H1N[⋯]O1ⁱ, hydrogen bond *cf.* Table 1, the dimers are connected further to form the two-dimensional layers.

Additionally, individual molecules of **1** interact with each other by means of π interactions. They are approximately arranged in a parallel-displaced sandwich type configuration and thus form one-dimensional layers along the crystallographic a -axis, *cf.* Figure 3. Within such a one-dimensional chain, a head-to-tail arrangement is observed, as reported previously for $\mathbf{1}\times 3\text{H}_2\text{O}$ (Oxtoby *et al.*, 2005). Moreover, by analogy with $\mathbf{1}\times 3\text{H}_2\text{O}$, the π stacks of **1** are arranged orthogonal to the two-dimensional layers formed by intermolecular hydrogen bonds. The combinations of both supramolecular arrangements finally give rise to a three-dimensional network structure.

S2. Experimental

Diethyl N,N' -naphthalene-bis(oxamate) was synthesized according to the literature (Rüffer *et al.*, 2007). To a solution of diethyl naphthalene-bis(oxamate) (1.5 g, 9.49 mmol) in MeOH (50 ml) three equivalents of a solution of *tert*-BuNH₂ (2.05 g, 28.47 mmol) in MeOH (25 ml) were added. The solution was refluxed for 30 minutes, cooled to room temperature and concentrated to 30 ml. Diethyl ether (100 ml) was added and the resulting brown precipitate was filtered and dried on air. Yellow crystals of **1** were obtained by solvent diffusion of a dilute MeOH solution of **1** against Et₂O at room temperature. Yield: 0.7 g, 80%.

S3. Refinement

C-bonded H atoms were placed in calculated positions and constrained to ride on their parent atoms, with a C—H distance of 0.93 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$. The N-bonded H atoms were located on a difference Fourier map and refined freely. The high R factor, low ratio of observed to unique reflections and relatively high s_u values indicate that the crystals were not of good quality and were very weakly diffracting.

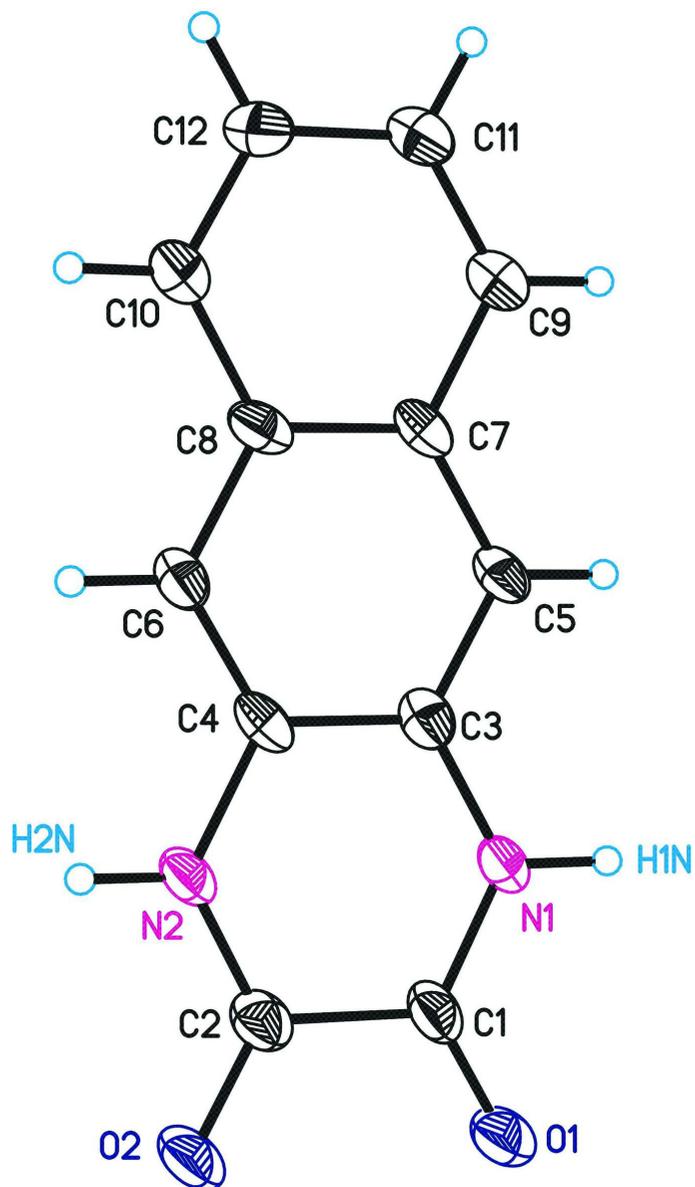
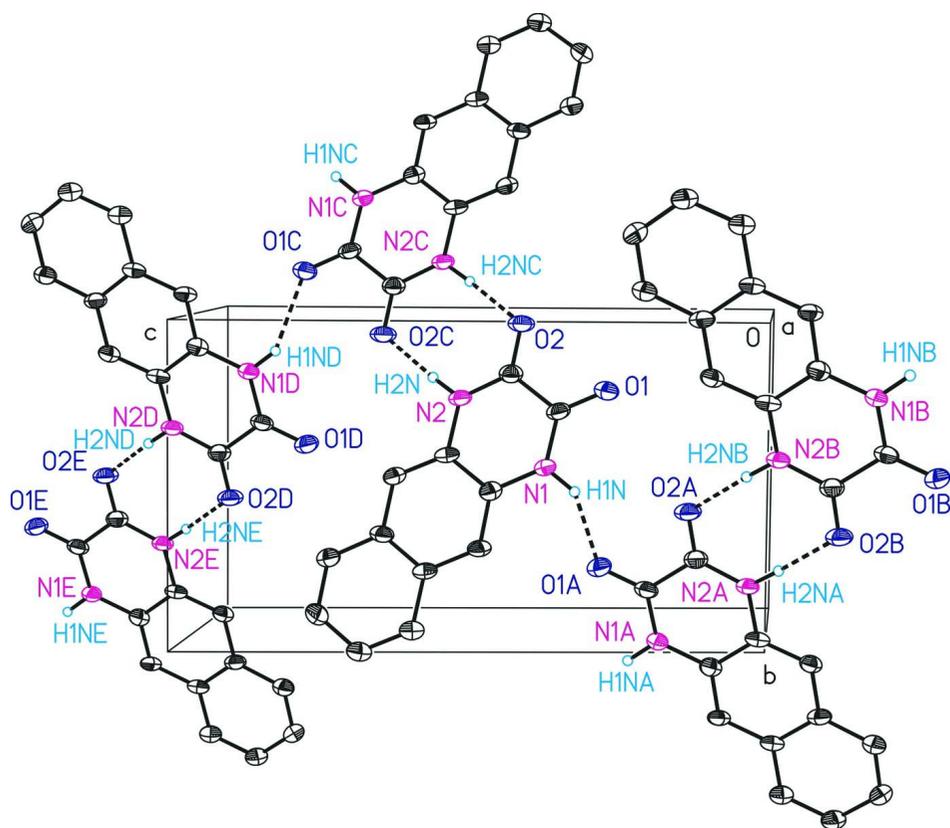
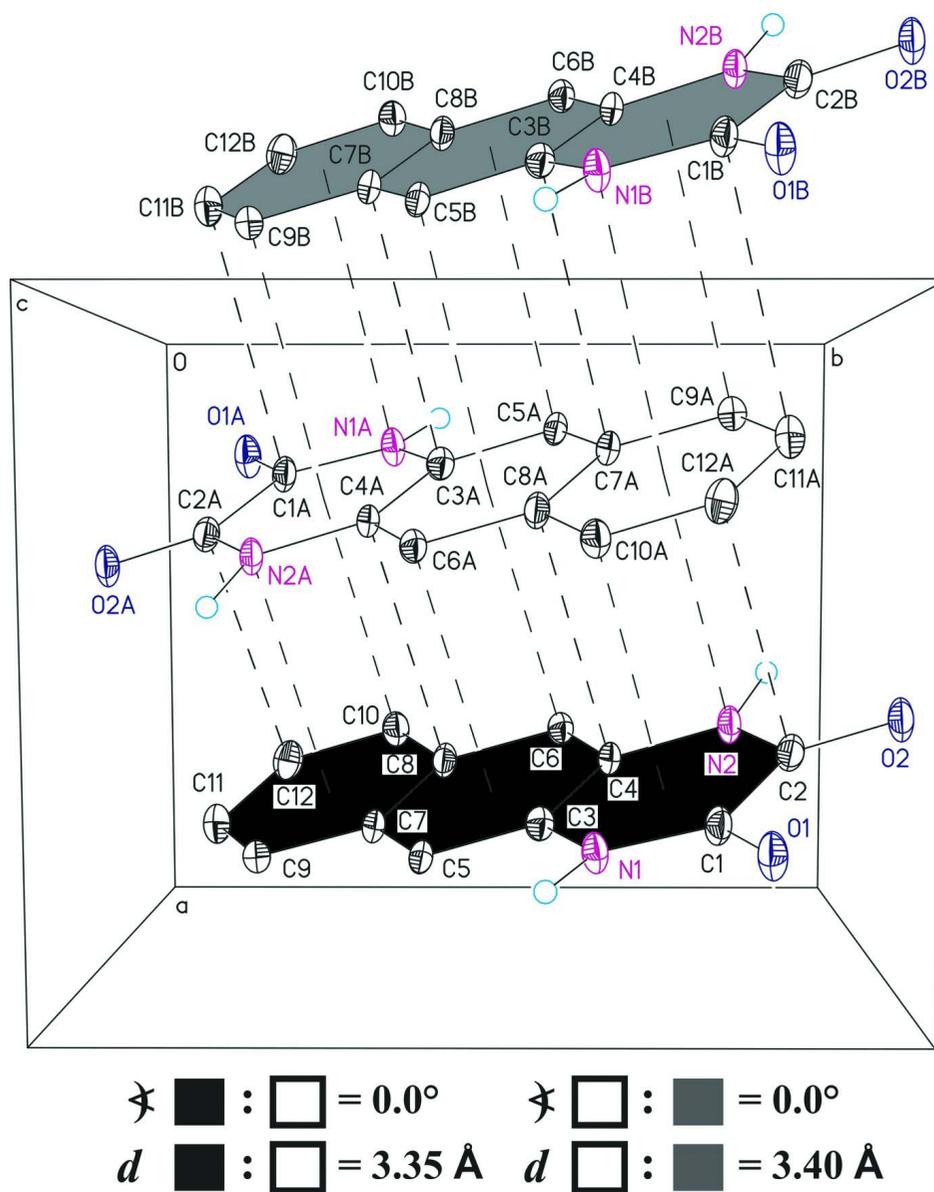


Figure 1

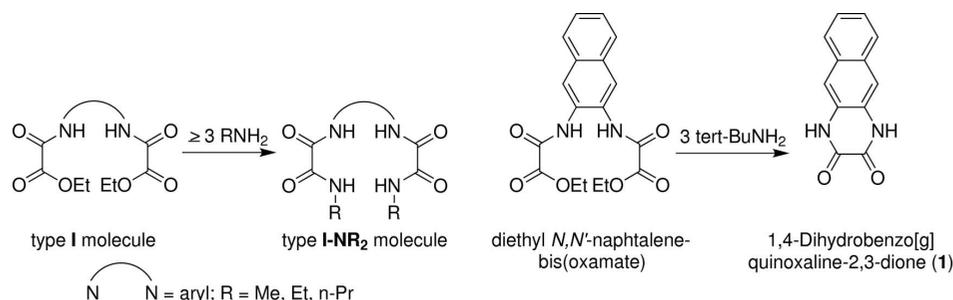
ORTEP diagram (50% ellipsoid probability) of the molecular structure of **1**.

**Figure 2**

Graphical representation of a part of one two-dimensional layer formed by **1** in the solid state due to formation of intermolecular hydrogen bonds along the crystallographic *b*- and *c*-axes. All C-bonded hydrogen atoms are omitted for clarity. Symmetry codes: (A) $-3/2 - x, 1/2 + y, 1/2 - z$. (B) $-1/2 + x, 1/2 - y, -1/2 + z$. (C) $-1 - x, -y, 1 - z$. (D) $-1/2 + x, 1/2 - y, 1/2 + z$. (E) $-1/2 - x, 1/2 + y, 3/2 - z$.


Figure 3

Graphical representation of the parallel-displaced sandwich type π interactions of **1** in the solid state, giving rise to the formation of one-dimensional chains along the crystallographic a -axes. All C-bonded hydrogen atoms are omitted for clarity. The sign \sphericalangle refers to the interplanar angle and d gives the averaged distances between interacting molecules. Dotted lines indicate interacting $C\cdots C$ atoms as well as interactions between carbon and nitrogen atoms, respectively, with the centroids of the respective adjacent aromatic rings. Symmetry codes: (A) $1 - x, 1 - y, 1 - z$. (B) $-1 + x, y, z$.


Figure 4

 Synthesis of **1**.

1,4-Dihydrobenzo[*g*]quinoxaline-2,3-dione
Crystal data
 $C_{12}H_8N_2O_2$
 $M_r = 212.20$

 Monoclinic, $P2_1/n$

Hall symbol: -P 2yn

 $a = 7.1334 (15) \text{ \AA}$
 $b = 8.4229 (18) \text{ \AA}$
 $c = 15.292 (2) \text{ \AA}$
 $\beta = 99.792 (14)^\circ$
 $V = 905.4 (3) \text{ \AA}^3$
 $Z = 4$
Data collection

Oxford Gemini S

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

 ω scans

Absorption correction: multi-scan

 (*CrysAlis RED*; Oxford Diffraction, 2006)

 $T_{\min} = 0.649$, $T_{\max} = 1.000$

5294 measured reflections

Refinement

 Refinement on F^2

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.094$
 $wR(F^2) = 0.271$
 $S = 0.92$

1773 reflections

153 parameters

0 restraints

Primary atom site location: structure-invariant

direct methods

 $F(000) = 440$
 $D_x = 1.557 \text{ Mg m}^{-3}$

 Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 5135 reflections

 $\theta = 2.9\text{--}25.9^\circ$
 $\mu = 0.11 \text{ mm}^{-1}$
 $T = 293 \text{ K}$

Block, yellow

 $0.3 \times 0.2 \times 0.1 \text{ mm}$

1773 independent reflections

 875 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.039$
 $\theta_{\max} = 26.0^\circ$, $\theta_{\min} = 3.0^\circ$
 $h = -8 \rightarrow 7$
 $k = -10 \rightarrow 8$
 $l = -18 \rightarrow 18$

2 standard reflections every 50 reflections

intensity decay: none

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H atoms treated by a mixture of independent and constrained refinement

 $w = 1/[\sigma^2(F_o^2) + (0.2P)^2]$

 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.77 \text{ e \AA}^{-3}$
 $\Delta\rho_{\min} = -0.57 \text{ e \AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.2609 (4)	-0.2869 (4)	1.1809 (2)	0.0276 (8)
C2	0.1631 (4)	-0.3842 (4)	1.1026 (2)	0.0276 (8)
C3	0.2764 (4)	-0.0662 (4)	1.07897 (19)	0.0257 (8)
C4	0.1859 (4)	-0.1580 (3)	1.0060 (2)	0.0247 (8)
C5	0.3312 (4)	0.0874 (4)	1.06546 (19)	0.0246 (8)
H5	0.3908	0.1472	1.1134	0.030*
C6	0.1489 (4)	-0.0948 (4)	0.92275 (19)	0.0245 (8)
H6	0.0876	-0.1557	0.8757	0.029*
C7	0.2984 (4)	0.1545 (4)	0.9806 (2)	0.0253 (8)
C8	0.2032 (4)	0.0626 (4)	0.9075 (2)	0.0262 (8)
C9	0.3508 (4)	0.3129 (4)	0.9647 (2)	0.0272 (8)
H9	0.4124	0.3738	1.0116	0.033*
C10	0.1628 (4)	0.1342 (4)	0.8220 (2)	0.0284 (8)
H10	0.0984	0.0763	0.7743	0.034*
C11	0.3125 (5)	0.3775 (4)	0.8816 (2)	0.0327 (8)
H11	0.3488	0.4815	0.8726	0.039*
C12	0.2174 (5)	0.2868 (4)	0.8089 (2)	0.0327 (9)
H12	0.1924	0.3313	0.7525	0.039*
N1	0.3053 (4)	-0.1348 (3)	1.16349 (17)	0.0273 (7)
N2	0.1300 (4)	-0.3136 (3)	1.02231 (17)	0.0260 (7)
O1	0.2918 (3)	-0.3464 (3)	1.25533 (14)	0.0361 (7)
O2	0.1113 (3)	-0.5216 (3)	1.11523 (14)	0.0344 (7)
H1N	0.354 (5)	-0.073 (5)	1.211 (2)	0.043 (10)*
H2N	0.053 (6)	-0.366 (5)	0.975 (3)	0.059 (12)*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0346 (19)	0.0156 (17)	0.0319 (17)	0.0018 (13)	0.0036 (13)	-0.0041 (13)
C2	0.0318 (18)	0.0181 (17)	0.0319 (16)	0.0028 (13)	0.0028 (13)	0.0005 (13)
C3	0.0279 (18)	0.0199 (18)	0.0290 (16)	0.0010 (12)	0.0041 (12)	-0.0026 (13)
C4	0.0258 (18)	0.0128 (16)	0.0358 (17)	0.0013 (12)	0.0055 (13)	-0.0030 (12)
C5	0.0273 (17)	0.0139 (16)	0.0316 (16)	0.0028 (12)	0.0023 (12)	-0.0012 (12)
C6	0.0265 (17)	0.0172 (17)	0.0289 (15)	0.0001 (12)	0.0019 (12)	-0.0036 (12)
C7	0.0279 (18)	0.0137 (17)	0.0345 (16)	0.0027 (12)	0.0056 (13)	-0.0020 (13)
C8	0.0307 (18)	0.0146 (17)	0.0347 (17)	0.0018 (12)	0.0099 (13)	0.0023 (13)

C9	0.0272 (18)	0.0195 (18)	0.0347 (17)	-0.0006 (13)	0.0044 (13)	-0.0005 (13)
C10	0.0333 (19)	0.0197 (18)	0.0322 (16)	-0.0007 (13)	0.0052 (13)	-0.0029 (13)
C11	0.042 (2)	0.0199 (18)	0.0370 (18)	-0.0007 (14)	0.0081 (14)	0.0021 (13)
C12	0.044 (2)	0.0243 (19)	0.0300 (16)	0.0043 (14)	0.0064 (14)	0.0065 (13)
N1	0.0394 (17)	0.0154 (15)	0.0257 (14)	-0.0026 (11)	0.0013 (11)	-0.0027 (10)
N2	0.0342 (16)	0.0130 (14)	0.0297 (14)	-0.0001 (10)	0.0021 (11)	-0.0018 (10)
O1	0.0532 (16)	0.0191 (14)	0.0330 (13)	-0.0021 (10)	-0.0007 (10)	0.0033 (9)
O2	0.0459 (15)	0.0148 (13)	0.0400 (14)	-0.0017 (9)	0.0005 (10)	0.0008 (9)

Geometric parameters (Å, °)

C1—O1	1.230 (4)	C6—H6	0.9300
C1—N1	1.356 (4)	C7—C9	1.417 (4)
C1—C2	1.518 (4)	C7—C8	1.432 (4)
C2—O2	1.240 (4)	C8—C10	1.425 (4)
C2—N2	1.349 (4)	C9—C11	1.366 (4)
C3—C5	1.378 (4)	C9—H9	0.9300
C3—N1	1.399 (4)	C10—C12	1.367 (4)
C3—C4	1.420 (4)	C10—H10	0.9300
C4—C6	1.364 (4)	C11—C12	1.423 (5)
C4—N2	1.404 (4)	C11—H11	0.9300
C5—C7	1.398 (4)	C12—H12	0.9300
C5—H5	0.9300	N1—H1N	0.91 (4)
C6—C8	1.412 (4)	N2—H2N	0.94 (4)
O1—C1—N1	123.8 (3)	C6—C8—C10	122.0 (3)
O1—C1—C2	119.7 (3)	C6—C8—C7	119.1 (3)
N1—C1—C2	116.5 (3)	C10—C8—C7	118.9 (3)
O2—C2—N2	122.8 (3)	C11—C9—C7	121.2 (3)
O2—C2—C1	119.4 (3)	C11—C9—H9	119.4
N2—C2—C1	117.7 (3)	C7—C9—H9	119.4
C5—C3—N1	121.7 (3)	C12—C10—C8	120.9 (3)
C5—C3—C4	119.8 (3)	C12—C10—H10	119.6
N1—C3—C4	118.5 (3)	C8—C10—H10	119.6
C6—C4—N2	121.0 (3)	C9—C11—C12	120.4 (3)
C6—C4—C3	120.7 (3)	C9—C11—H11	119.8
N2—C4—C3	118.3 (3)	C12—C11—H11	119.8
C3—C5—C7	120.8 (3)	C10—C12—C11	120.1 (3)
C3—C5—H5	119.6	C10—C12—H12	120.0
C7—C5—H5	119.6	C11—C12—H12	120.0
C4—C6—C8	120.4 (3)	C1—N1—C3	124.6 (3)
C4—C6—H6	119.8	C1—N1—H1N	117 (2)
C8—C6—H6	119.8	C3—N1—H1N	118 (2)
C5—C7—C9	122.2 (3)	C2—N2—C4	124.2 (3)
C5—C7—C8	119.2 (3)	C2—N2—H2N	119 (3)
C9—C7—C8	118.6 (3)	C4—N2—H2N	117 (3)
O1—C1—C2—O2	1.7 (5)	C5—C7—C8—C10	177.3 (3)

N1—C1—C2—O2	-176.8 (3)	C9—C7—C8—C10	-0.8 (4)
O1—C1—C2—N2	179.3 (3)	C5—C7—C9—C11	-178.1 (3)
N1—C1—C2—N2	0.8 (4)	C8—C7—C9—C11	-0.1 (5)
C5—C3—C4—C6	-1.3 (5)	C6—C8—C10—C12	-179.8 (3)
N1—C3—C4—C6	177.9 (3)	C7—C8—C10—C12	1.5 (5)
C5—C3—C4—N2	-179.3 (3)	C7—C9—C11—C12	0.4 (5)
N1—C3—C4—N2	-0.1 (4)	C8—C10—C12—C11	-1.3 (5)
N1—C3—C5—C7	-179.0 (3)	C9—C11—C12—C10	0.4 (5)
C4—C3—C5—C7	0.2 (5)	O1—C1—N1—C3	178.1 (3)
N2—C4—C6—C8	179.0 (2)	C2—C1—N1—C3	-3.5 (4)
C3—C4—C6—C8	1.1 (5)	C5—C3—N1—C1	-177.6 (3)
C3—C5—C7—C9	179.2 (3)	C4—C3—N1—C1	3.2 (5)
C3—C5—C7—C8	1.2 (5)	O2—C2—N2—C4	179.7 (3)
C4—C6—C8—C10	-178.4 (3)	C1—C2—N2—C4	2.2 (4)
C4—C6—C8—C7	0.3 (5)	C6—C4—N2—C2	179.5 (3)
C5—C7—C8—C6	-1.4 (4)	C3—C4—N2—C2	-2.6 (5)
C9—C7—C8—C6	-179.5 (3)		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1N...O1 ⁱ	0.91 (4)	2.27 (4)	2.866 (3)	122 (3)
N2—H2N...O2 ⁱⁱ	0.94 (4)	1.90 (4)	2.843 (4)	176 (4)

Symmetry codes: (i) $-x+1/2, y+1/2, -z+5/2$; (ii) $-x, -y-1, -z+2$.