

Dihydrionium tetrachromate(VI), $(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$

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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{Cr}=\text{O}) = 0.005$ Å; H-atom completeness 0%; R factor = 0.040; wR factor = 0.110; data-to-parameter ratio = 15.5.

The crystal structure of $(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$ is isotropic with $\text{K}_2\text{Cr}_4\text{O}_{13}$. The finite tetrachromate anion in the title structure consists of four vertex-sharing CrO_4 tetrahedra and exhibits a typical zigzag arrangement. The crystal packing is stabilized by hydrogen bonds between these anions and hydronium cations. The two different hydronium cations are surrounded by nine O atoms of tetrachromate anions, with $\text{O}\cdots\text{O}$ distances ranging between 2.866 (8) and 3.282 (7) Å.

Related literature

The title chromate is isotropic with its potassium analogue (Casari & Langer, 2005). Löfgren (1973) and Kolitsch (2004) determined the structures of the corresponding Rb and Cs salts, respectively. For industrial applications of tetrachromates, see: Cainelli & Cardillo (1984); Çengeloglu *et al.* (2003). For related bond-length data, see: Casari *et al.* (2007). For cell parameters of further isolated compounds stated in the experimental procedure, see: Durif & Averbuch-Pouchot (1978) and Rahman *et al.* (2003).

Experimental

Crystal data

$(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$
 $M_r = 454.05$
Monoclinic, Pc

$a = 8.9765$ (13) Å
 $b = 7.6431$ (8) Å
 $c = 9.3451$ (14) Å

$\beta = 91.888$ (18)°
 $V = 640.80$ (15) Å³
 $Z = 2$
Mo $K\alpha$ radiation

$\mu = 3.37$ mm⁻¹
 $T = 293$ K
 $1.0 \times 0.4 \times 0.2$ mm

Data collection

Stoe IPDS I diffractometer
Absorption correction: numerical
[X-RED (Stoe & Cie, 2001) and
X-SHAPE (Stoe & Cie, 1999)]
 $T_{\min} = 0.121$, $T_{\max} = 0.314$

5900 measured reflections
2696 independent reflections
2497 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.040$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.040$
 $wR(F^2) = 0.110$
 $S = 1.06$
2696 reflections
174 parameters
2 restraints

$\Delta\rho_{\text{max}} = 0.69$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.58$ e Å⁻³
Absolute structure: Flack (1983),
1212 Friedel pairs
Flack parameter: 0.53 (4)

Data collection: IPDS (Stoe & Cie, 1997); cell refinement: IPDS; data reduction: IPDS; program(s) used to solve structure: SIR92 (Altomare *et al.*, 1993); program(s) used to refine structure: SHEXL97 (Sheldrick, 2008); molecular graphics: DIAMOND (Crystal Impact, 2012); software used to prepare material for publication: SHEXL97.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WM2714).

References

- Altomare, A., Casciaro, G., Giacovazzo, C. & Guagliardi, A. (1993). *J. Appl. Cryst.* **26**, 343–350.
- Cainelli, G. & Cardillo, G. (1984). In *Chromium Oxidations in Organic Chemistry*. Berlin: Springer.
- Casari, B. M., Eriksson, A. K. & Langer, V. (2007). *Z. Naturforsch. Teil B*, **62**, 771–777.
- Casari, B. M. & Langer, V. (2005). *Acta Cryst. C* **61**, i117–i119.
- Çengeloglu, Y., Tor, A., Kir, E. & Ersöz, M. (2003). *Desalination*, **154**, 239–246.
- Crystal Impact (2012). DIAMOND. Crystal Impact GbR, Bonn, Germany.
- Durif, A. & Averbuch-Pouchot, M. T. (1978). *Acta Cryst. B* **34**, 3335–3337.
- Flack, H. D. (1983). *Acta Cryst. A* **39**, 876–881.
- Kolitsch, U. (2004). *Acta Cryst. C* **60**, i17–i19.
- Löfgren, P. (1973). *Acta Cryst. B* **29**, 2141–2147.
- Rahman, A. A., Usman, A., Chantrapromma, S. & Fun, H.-K. (2003). *Acta Cryst. C* **59**, i92–i94.
- Sheldrick, G. M. (2008). *Acta Cryst. A* **64**, 112–122.
- Stoe & Cie (1997). IPDS. Stoe & Cie, Darmstadt, Germany.
- Stoe & Cie (1999). X-SHAPE. Stoe & Cie, Darmstadt, Germany.
- Stoe & Cie (2001). X-RED. Stoe & Cie, Darmstadt, Germany.

supporting information

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Dihydrionium tetrachromate(VI), $(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$

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S1. Comment

One of the characteristic properties of the elements of group 6 of the periodic table is their ability to build anionic polyoxo compounds with the corresponding elements in high oxidation states - the so called polyoxometalates. This property is profoundly exhibited by the oxoanions of the heavier elements Mo(VI) and W(VI), whereas the respective compounds of Cr(VI), i.e. polyoxochromates, are not built readily. Despite of this fact, polyoxochromates are of chemical and industrial importance due to their high oxidation potential. Accordingly, they are used as oxidants of organic compounds (Cainelli & Cardillo, 1984) and in hexavalent chromium plating (Çengeloğlu *et al.*, 2003).

The title compound, $(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$, is only the fourth tetrachromate(VI) described and characterized by single-crystal X-ray diffraction so far. The first three are salts of alkali metals, *viz.* the potassium (Casari & Langer, 2005), the rubidium (Löfgren, 1973), and the caesium salt (Kolitsch, 2004). $(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$ was isolated from a reaction mixture containing $\text{Na}_2\text{Cr}_2\text{O}_7$, nitric acid and a large excess of CrO_3 along with several possibly unrelated species.

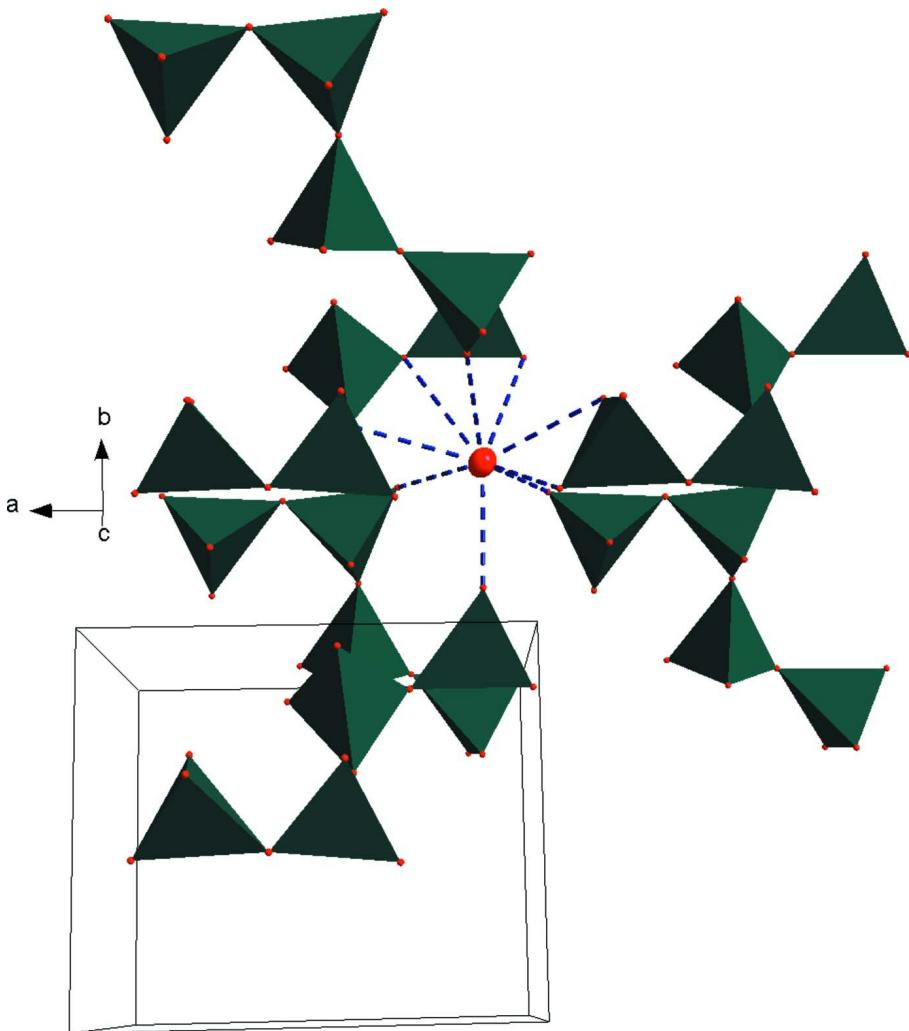
The crystal structure of $(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$ is isotypic with that of $\text{K}_2\text{Cr}_4\text{O}_{13}$ (Casari & Langer, 2005) and accordingly exhibits the space group *Pc*. The cell volume of $(\text{H}_3\text{O})_2(\text{Cr}_4\text{O}_{13})$ is $640.80(15)$ Å³ at room temperature. It exceeds the cell volume of the potassium analogue measured at 173 K by roughly 44 Å³. The finite tetrachromate anion is composed of four condensed CrO_4 tetrahedra and exhibits the typical zigzag arrangement (Figs. 1,2) described by Casari & Langer (2005). The two hydronium ions have two crystallographically different positions. They interact with nine oxygen atoms of the tetrachromate anions through hydrogen bonds. Although the H atoms of the hydronium cations could not be located, the distances between the hydronium O atoms and the surrounding tetrachromate O atoms between 2.866 (8) and 3.282 (7) Å point to moderate to weak O—H···O hydrogen bonds. These distances are quite similar to the distances between O atoms of water molecules and dichromate ions in $\text{Na}_2\text{Cr}_2\text{O}_7\text{H}_2\text{O}$ (Casari *et al.*, 2007) which indicates hydrogen bonds of typical strength for this class of compounds.

S2. Experimental

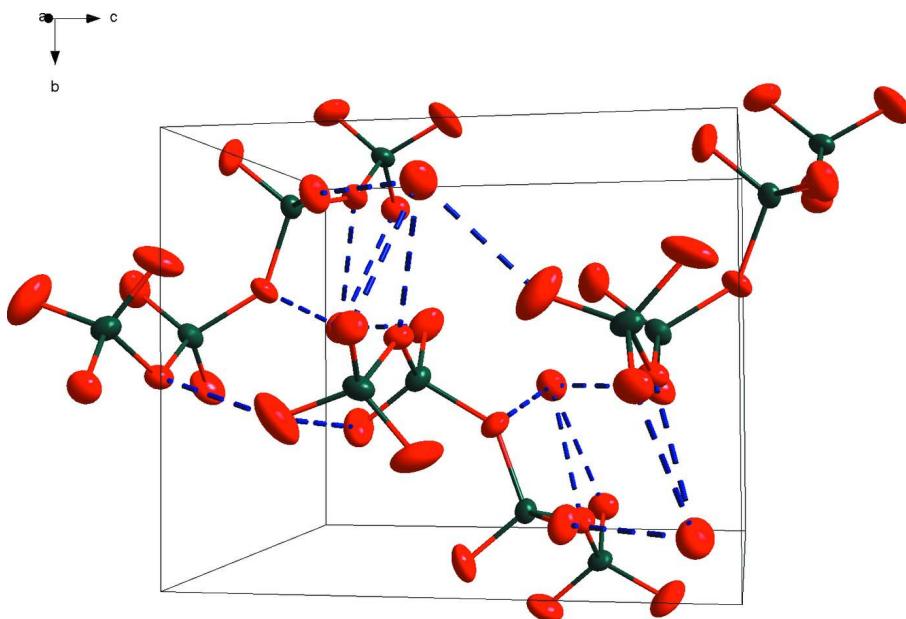
$\text{Na}_2\text{Cr}_2\text{O}_7$ (26 mg, 0.1 mmol) and CrO_3 (1.600 g, 16 mmol) were dissolved in 0.5 ml H_2O . This solution was added to a solution of AgClO_4 (22.5 mg, 0.1 mmol) and theobromine (18 mg, 0.1 mmol) in 16.5 ml of nitric acid. After 2.5 months crystals of $(\text{H}_3\text{O})(\text{ClO}_4)$ (Rahman *et al.*, 2003) and $\text{Ag}_2(\text{Cr}_2\text{O}_7)$ (Durif & Averbuch-Pouchot, 1978) were isolated and characterized by X-ray diffractometric unit-cell determinations. Orange-red crystals of the title compound were obtained from the mother liquor after another half year.

S3. Refinement

The investigated crystal was racemically twinned, similarly to the potassium compound (Casari & Langer, 2005). The refined Flack parameter indicates a twin component ratio of 53 (4):47 (4). It was not possible to unambiguously locate the H atoms of the hydronium cations. They were therefore omitted from the refinement.

**Figure 1**

Hydrogen bonds between the hydronium ion ($\text{O}1\text{W}$) and the surrounding tetrachromate anions. The oxygen atom of the hydronium ion is shown with anisotropic displacement parameters at the 50% probability level. Chromate anions are shown as green-blue tetrahedra. Dashed lines denote $\text{O}\cdots\text{O}$ contacts between the cation and the anions.

**Figure 2**

Hydrogen bonding framework within the unit cell of the compound. See Fig. 1 for legend.

Dihydronium tetrachromate(VI)

Crystal data

$(\text{H}_3\text{O})_2\text{Cr}_4\text{O}_{13}$
 $M_r = 454.05$
Monoclinic, Pc
Hall symbol: P -2yc
 $a = 8.9765 (13) \text{ \AA}$
 $b = 7.6431 (8) \text{ \AA}$
 $c = 9.3451 (14) \text{ \AA}$
 $\beta = 91.888 (18)^\circ$
 $V = 640.80 (15) \text{ \AA}^3$
 $Z = 2$

Data collection

Stoe IPDS I
diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
 φ scans
Absorption correction: numerical
[X-RED (Stoe & Cie, 2001) and X-SHAPE (Stoe & Cie, 1999)]
 $T_{\min} = 0.121$, $T_{\max} = 0.314$

$F(000) = 444$
 $D_x = 2.353 \text{ Mg m}^{-3}$
Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
Cell parameters from 1754 reflections
 $\theta = 1.9\text{--}28.2^\circ$
 $\mu = 3.37 \text{ mm}^{-1}$
 $T = 293 \text{ K}$
Block, orange-red
 $1.0 \times 0.4 \times 0.2 \text{ mm}$

Refinement

Refinement on F^2
Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.040$
 $wR(F^2) = 0.110$
 $S = 1.06$
2696 reflections
174 parameters
2 restraints
Primary atom site location: structure-invariant direct methods

Secondary atom site location: difference Fourier map
 $w = 1/[\sigma^2(F_o^2) + (0.0842P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 0.69 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.58 \text{ e } \text{\AA}^{-3}$

Extinction correction: *SHELXL97* (Sheldrick, 2008), $F_c^* = kFc[1 + 0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
 Extinction coefficient: 0.083 (5)
 Absolute structure: Flack (1983), 1212 Friedel pairs
 Absolute structure parameter: 0.53 (4)

Special details

Experimental. A suitable single-crystal was carefully selected under a microscope and mounted in a glass capillary. The scattering intensities were collected on an imaging plate diffractometer (*IPDS I*, Stoe & Cie) equipped with a fine focus sealed tube X-ray source (Mo K α , $\lambda = 0.71073 \text{ \AA}$) operating at 50 kV and 40 mA. Intensity data for the title compound were collected at room temperature by φ -scans in 100 frames ($0 < \varphi < 200^\circ$, $\Delta\varphi = 2^\circ$, exposure time of 7 min) in the 2Θ range 3.8 to 56.3° . Structure solution and refinement were carried out using the programs *SIR92* (Altomare *et al.*, 1993) and *SHELXL97* (Sheldrick, 1997) embedded into *WinGX* program package (Farrugia, 1999). A numerical absorption correction (*X-RED* (Stoe & Cie, 2001) was applied after optimization of the crystal shape (*X-SHAPE* (Stoe & Cie, 1999)). The last cycles of refinement included atomic positions and anisotropic parameters for all non-hydrogen atoms. Positions of hydrogen atoms were not determined. The final difference maps were free of any chemically significant features. The refinement was based on F^2 for ALL reflections.

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Cr2	0.44108 (9)	1.10692 (11)	1.05710 (8)	0.0219 (2)
Cr1	0.12169 (8)	0.92974 (12)	1.17627 (8)	0.0242 (2)
Cr3	0.44219 (9)	1.44260 (12)	0.83187 (8)	0.0263 (2)
Cr4	0.78483 (9)	1.42919 (12)	0.79827 (9)	0.0283 (2)
O21	0.4560 (6)	0.9751 (8)	0.9295 (4)	0.0437 (12)
O31	0.4312 (7)	1.3165 (9)	0.7003 (5)	0.0542 (13)
O14	0.2882 (5)	1.0642 (5)	1.1542 (4)	0.0324 (9)
O33	0.6117 (5)	1.5476 (5)	0.8324 (4)	0.0296 (8)
O12	0.1024 (6)	0.7916 (8)	1.0473 (6)	0.0497 (12)
O22	0.5891 (5)	1.0957 (7)	1.1545 (4)	0.0401 (11)
O23	0.4224 (5)	1.3239 (6)	0.9907 (4)	0.0334 (9)
O43	0.9211 (5)	1.5638 (6)	0.8196 (5)	0.0360 (10)
O11	-0.0193 (5)	1.0612 (6)	1.1705 (5)	0.0395 (11)
O13	0.1342 (5)	0.8307 (8)	1.3278 (6)	0.0522 (14)
O42	0.7792 (7)	1.3587 (12)	0.6374 (7)	0.074 (2)
O32	0.3119 (5)	1.5822 (7)	0.8197 (6)	0.0477 (13)
O41	0.8015 (6)	1.2744 (8)	0.9121 (7)	0.0595 (16)
O1W	0.0967 (6)	0.4158 (8)	1.0652 (6)	0.0469 (12)
O2W	0.7959 (6)	0.9005 (7)	0.9169 (5)	0.0456 (12)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cr2	0.0205 (3)	0.0256 (4)	0.0197 (3)	-0.0006 (3)	0.0010 (2)	0.0020 (3)
Cr1	0.0221 (4)	0.0208 (5)	0.0296 (4)	-0.0019 (3)	0.0017 (3)	0.0052 (3)
Cr3	0.0235 (4)	0.0274 (5)	0.0278 (4)	-0.0021 (3)	0.0004 (3)	0.0068 (3)
Cr4	0.0255 (4)	0.0222 (5)	0.0374 (4)	-0.0005 (3)	0.0057 (3)	-0.0052 (3)
O21	0.047 (3)	0.060 (4)	0.0242 (18)	0.010 (2)	0.0025 (17)	-0.0114 (18)
O31	0.063 (3)	0.062 (4)	0.038 (2)	-0.024 (3)	0.0033 (19)	-0.008 (2)
O14	0.034 (2)	0.029 (3)	0.0349 (19)	-0.0062 (15)	0.0083 (16)	0.0006 (13)
O33	0.0289 (19)	0.024 (2)	0.036 (2)	-0.0003 (16)	0.0045 (15)	0.0024 (14)
O12	0.048 (3)	0.035 (3)	0.065 (3)	-0.003 (2)	-0.004 (2)	-0.017 (2)
O22	0.027 (2)	0.058 (3)	0.035 (2)	-0.0030 (18)	-0.0041 (16)	0.0082 (17)
O23	0.038 (2)	0.028 (2)	0.0344 (19)	-0.0032 (18)	0.0057 (15)	0.0107 (14)
O43	0.026 (2)	0.036 (3)	0.046 (2)	-0.0044 (15)	0.0092 (17)	0.0016 (16)
O11	0.030 (2)	0.040 (3)	0.049 (2)	0.0044 (16)	0.0039 (18)	0.0018 (17)
O13	0.038 (2)	0.061 (4)	0.057 (3)	-0.010 (2)	0.002 (2)	0.034 (3)
O42	0.059 (4)	0.092 (5)	0.071 (4)	0.006 (3)	0.007 (3)	-0.047 (4)
O32	0.029 (3)	0.057 (4)	0.057 (3)	0.006 (2)	-0.0039 (19)	0.022 (2)
O41	0.048 (3)	0.034 (3)	0.097 (4)	0.002 (2)	0.010 (3)	0.026 (3)
O1W	0.044 (3)	0.051 (3)	0.046 (2)	0.007 (2)	0.0014 (19)	-0.0030 (19)
O2W	0.043 (3)	0.049 (4)	0.044 (2)	0.009 (2)	-0.0039 (19)	0.0003 (19)

Geometric parameters (\AA , $^\circ$)

Cr2—O21	1.570 (4)	Cr3—O31	1.563 (6)
Cr2—O22	1.588 (4)	Cr3—O32	1.584 (5)
Cr2—O14	1.701 (4)	Cr3—O33	1.720 (5)
Cr2—O23	1.776 (4)	Cr3—O23	1.754 (4)
Cr1—O13	1.606 (4)	Cr4—O41	1.595 (5)
Cr1—O12	1.607 (5)	Cr4—O42	1.596 (6)
Cr1—O11	1.615 (5)	Cr4—O43	1.606 (5)
Cr1—O14	1.831 (4)	Cr4—O33	1.836 (5)
O21—Cr2—O22	108.1 (3)	O32—Cr3—O33	109.7 (3)
O21—Cr2—O14	111.9 (2)	O31—Cr3—O23	110.0 (3)
O22—Cr2—O14	111.0 (2)	O32—Cr3—O23	108.3 (2)
O21—Cr2—O23	110.1 (3)	O33—Cr3—O23	110.7 (2)
O22—Cr2—O23	108.5 (2)	O41—Cr4—O42	112.2 (4)
O14—Cr2—O23	107.3 (2)	O41—Cr4—O43	109.7 (3)
O13—Cr1—O12	110.7 (3)	O42—Cr4—O43	109.5 (4)
O13—Cr1—O11	110.8 (3)	O41—Cr4—O33	108.1 (2)
O12—Cr1—O11	108.6 (3)	O42—Cr4—O33	109.3 (3)
O13—Cr1—O14	109.3 (2)	O43—Cr4—O33	108.0 (2)
O12—Cr1—O14	110.6 (2)	Cr2—O14—Cr1	147.5 (3)
O11—Cr1—O14	106.8 (2)	Cr3—O33—Cr4	121.6 (2)
O31—Cr3—O32	109.3 (3)	Cr3—O23—Cr2	140.1 (3)
O31—Cr3—O33	108.9 (3)		